SMART TEST SERIES

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Name:		Subject:	Chemistry-12
Roll # :		Unit(s):	1,
Class:	Inter Part-II	Test:	Type 3 - MCQs + SQs Test - Marks=30
Date:		Time:	

Q.1 Four possible answers A, B, C & D to each question are given. Circle the correct one. (10x1=10)

1 In modern periodic table 6th period contains elements?
(A) 8 (B) 18 (C) 10 (D) 32

2 Which is the longest period of periodic table:-
(A) 4 (B) 5 (C) 6 (D) 7

3 Which of the following has highest M.P?
(A) Aluminium (B) Silicon (C) Phosphorus (D) Sulphur

4 Which of the following has the highest hydration energy?
(A) Li^+ (B) Na^+ (C) K^+ (D) Mg^{++}

5 The element of 2nd period, which has highest ionization energy from the following is:-
(A) Be (B) C (C) N (D) O

6 Which one is not a periodic property:
(A) Ionization energy (B) Density (C) Atomic Radii (D) Hydration energy

7 Hydrogen resembles in properties with groups:
(A) I-A, V-A, VII-A elements (B) I-A, IV-A, VII-A elements (C) II-A, III-A, V-A elements
(D) I-A, II-A, elements

8 Mark the correct statement:
(A) All lanthanides are present in the same group.
(B) All halogens are present in the same period.
(C) All the alkali metals are present in the same group.
(D) All the noble gases are present in the same period.

9 Mark the correct statement:
(A) Metallic character increases down the group. (B) Metallic character increases along a period.
(C) Metallic character decreases along a period.
(D) Metallic character remains the same down the group

10 Mark the correct statement:
(A) Melting points of halogens decrease down the group.
(B) Melting points of halogens increase down the group.
(C) Melting points of halogens remain the same
(D) Melting points of halogens first increase and then decrease down the group.

Q.2 Write short answers of the following questions. (10x2=20)

(i) What is Dobriener's law of triads?
(ii) M.P,B.P of short periods increases upto middle of the period and then decrease. Why?
(iii) Define atomic radius. Why atomic radius of IA group elements increases in a group?
(iv) Explain the variation in melting points along the short periods.
(v) How does lanthanide contraction control the atomic size of elements of 6th and 7th periods?
(vi) What are amphoteric oxides? Give an example?
(vii) What are Ionic hydrides? Give example.
(viii) Define Covalent Hydride with one example.
(ix) Give any two resemblances of hydrogen with Group IV-A
(x) Justify the position of hydrogen at the top of VII A group.