



SMART TEST SERIES

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Name:		Subject:	Chemistry-11
Roll # :		Unit(s):	3,
Class:	Inter Part-I	Test:	Type 1 - MCQs Test - Marks=20
Date:		Time:	

Q.1 Four possible answers A, B, C & D to each question are given. Circle the correct one.

(20x1=20)

- Borax has the chemical formula:
(A) KNO_3 (B) $NaNO_3$ (C) $Na_2B_4O_7 \cdot 10 H_2O$ (D) $Na_2CO_3 \cdot H_2O$
- If absolute temperature of the gas is doubled and the pressure is reduced to one half the volume of the gas will:
(A) Remains unchanged (B) Increase four times (C) Reduce to $\frac{1}{4}$ (D) be doubled
- Pressure remaining constant at which temperature the volume of a gas will become twice of what it is at $0^\circ C$:
(A) $546^\circ C$ (B) $200^\circ C$ (C) $546 K$ (D) $273 K$
- The unit millibar is commonly used by:
(A) Meteorologists (B) Astronauts (C) Engineers (D) Dalton
- The number of molecules in one dm^3 of water is close to:
(A) $\frac{6.02}{22.4} \times 10^{23}$ (B) $\frac{12.04}{22.4} \times 10^{23}$ (C) $\frac{18}{22.4} \times 10^{23}$ (D) $55.6 \times 6.02 \times 10^{23}$
- Equal masses of methane and oxygen are mixed in an empty container at $25^\circ C$. The fraction of total pressure exerted by oxygen is:
(A) $\frac{1}{3}$ (B) $\frac{8}{9}$ (C) $\frac{1}{9}$ (D) $\frac{16}{17}$
- Pressure remaining constant at which temperature the volume of a gas will become twice of what it is at $0^\circ C$:
(A) $546^\circ C$ (B) $200^\circ C$ (C) $546 K$ (D) $273 K$
- Vapour pressure of liquid depends upon:
(A) Amount of liquid (B) Surface area (C) Temperature (D) Size of container
- The S.I unit of pressure is:
(A) Torr (B) mmHg (C) Pounds $inch^{-2}$ (D) Nm^{-2}
- The commonly used unit of pressure by meteorologists is:
(A) Atmosphere (B) Pascal (C) Millibar (D) Pound $inch^{-2}$
- Feeling uncomfortable breathing in un-pressurized cabins is due to:
(A) High pressure of CO_2 (B) Low Pressure of O_2 (C) Fatigue (D) Low pressure of CO_2
- The order of the rate of diffusion of gases NH_3 , SO_2 , Cl_2 and CO_2 is:
(A) $NH_3 > SO_2 > Cl_2 > CO_2$ (B) $NH_3 > CO_2 > SO_2 > Cl_2$
(C) $Cl_2 > SO_2 > CO_2 > NH_3$ (D) $NH_3 > CO_2 > Cl_2 > SO_2$
- Which of the following will have highest rate of diffusion?
(A) O_2 (B) CO_2 (C) NH_3 (D) SO_2
- The highest temperature at which a substance can exist as liquid at its critical pressure is:
(A) Absolute zero (B) Consulate temperature (C) Critical temperature (D) Transition temperature
- The temperature of natural plasma is about:
(A) $20000^\circ C$ (B) $10000^\circ C$ (C) $5000^\circ C$ (D) $1000^\circ C$
- Which of the following will have the same number of molecules at STP?
(A) $280 cm^3$ of CO_2 and $280 cm^3$ of N_2O (B) $11.2 dm^3$ of O_2 and $32 g$ of O_2
(C) $44 g$ of CO_2 and $11.2 dm^3$ of CO (D) $28 g$ of N_2 and $5.6 dm^3$ of oxygen
- How should the conditions be changed to prevent the volume of a give gas from expanding when its mass is increased?
(A) Temperature is lowered and pressure is increased
(B) Temperature is increased and pressure is lowered (C) Temperature and pressure both are lowered
(D) Temperature and pressure both are increased.
- The molar volume of CO_2 is maximum at:
(A) STP (0° and $1 atm$) (B) $127^\circ C$ and $1 atm$ (C) $0^\circ C$ and $2 atm$ (D) $273^\circ C$ and $2 atm$
- Equal masses of methane and oxygen are mixed in an empty container at $25^\circ C$. The fraction of total pressure exerted by oxygen is:
(A) $\frac{1}{3}$ (B) $\frac{8}{9}$ (C) $\frac{1}{9}$ (D) $\frac{16}{17}$
- The molar volume of CO_2 in maximum at:
(A) S.T.P (B) $127^\circ C$ and $1 atm$ (C) $0^\circ C$ and $2 atm$ (D) $273^\circ C$ and $2 atm$

MCQs Ans Key.

Q:1 (C)

Q:2 (B)

Q:3 (C)

Q:4 (A)

Q:5 (D)

Q:6 (A)

Q:7 (C)

Q:8 (C)

Q:9 (D)

Q:10 (B)

Q:11 (B)

Q:12 (B)

Q:13 (C)

Q:14 (C)

Q:15 (A)

Q:16 (A)

Q:17 (A)

Q:18 (B)

Q:19 (A)

Q:20 (B)