



SMART TEST SERIES

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|----------|--------------|----------|-------------------------------|
| Name: | | Subject: | Chemistry-11 |
| Roll # : | | Unit(s): | 3, |
| Class: | Inter Part-I | Test: | Type 1 - MCQs Test - Marks=20 |
| Date: | | Time: | |

Q.1 Four possible answers A, B, C & D to each question are given. Circle the correct one.

(20x1=20)

- If absolute temperature of the gas is doubled and the pressure is reduced to one half the volume of the gas will:
(A) Remains unchanged (B) Increase four times (C) Reduce to $\frac{1}{4}$ (D) be doubled
- Pressure remaining constant at which temperature the volume of a gas will become twice of water it is at 0°C :
(A) 546°C (B) 200°C (C) 546K (D) 273K
- Formula used for the conversion of F° into C° is:
(A) (B) $C^{\circ} = \frac{5}{9}[F^{\circ} - 32]$ (C) $F^{\circ} = \frac{5}{9}(C^{\circ}) + 32$ (D) $C^{\circ} = \frac{9}{5}[F^{\circ} - 32]$
- The unit millibar is commonly used by:
(A) Meteorologists (B) Astronauts (C) Engineers (D) Dalton
- The sum of mole fraction of gas in a mixture of gases is:
(A) Always more than one (B) Always less than one (C) Always one
(D) May be less or more than one
- The number of molecules in one dm^3 of water is close to:
(A) $\frac{6.02}{22.4} \times 10^{23}$ (B) $\frac{12.04}{22.4} \times 10^{23}$ (C) $\frac{18}{22.4} \times 10^{23}$ (D) $55.6 \times 6.02 \times 10^{23}$
- Partial pressure of oxygen in the air is:
(A) 156 torr (B) 157 torr (C) 158 torr (D) 159 torr
- Vapour pressure of liquid depends upon:
(A) Amount of liquid (B) Surface area (C) Temperature (D) Size of container
- The S.I unit of pressure is:
(A) Torr (B) mmHg (C) Pounds inch^{-2} (D) Nm^{-2}
- The commonly used unit of pressure by meteorologists is:
(A) Atmosphere (B) Pascal (C) Millibar (D) Pound inch^{-2}
- The partial pressure of oxygen in lungs is:
(A) 760 torr (B) 320 torr (C) 159 torr (D) 116 torr
- Feeling uncomfortable breathing in un-pressurized cabins is due to:
(A) High pressure of CO_2 (B) Low Pressure of O_2 (C) Fatigue (D) Low pressure of CO_2
- Feeling uncomfortable breathing in un-pressurized cabins is due to:
(A) High pressure of CO_2 (B) Low Pressure of O_2 (C) Fatigue (D) Low pressure of CO_2
- The order of the rate of diffusion of gases NH_3 , SO_2 , Cl_2 and CO_2 is:
(A) $\text{NH}_3 > \text{SO}_2 > \text{Cl}_2 > \text{CO}_2$ (B) $\text{NH}_3 > \text{CO}_2 > \text{SO}_2 > \text{Cl}_2$ (C) $\text{Cl}_2 > \text{SO}_2 > \text{CO}_2 > \text{NH}_3$
(D) $\text{NH}_3 > \text{CO}_2 > \text{Cl}_2 > \text{SO}_2$
- Which of the following will have highest rate of diffusion?
(A) O_2 (B) CO_2 (C) NH_3 (D) SO_2
- The deviation of a gas from ideal behaviour is maximum at:
(A) -10°C and 5.0 atm (B) -10°C and 2.0 (C) 100°C and 2 atm (D) 0°C and 2.0 atm
- A real gas obeying Van der Waal's equation will resemble ideal gas if:
(A) both a and b are large (B) both a and b are small or zero (C) a is small and b is large
(D) a is large and b is small
- The highest temperature at which a substance can exist as liquid at its critical pressure is:
(A) Absolute zero (B) Consulate temperature (C) Critical temperature (D) Transition temperature
- The temperature of natural plasma is about:
(A) 20000°C (B) 10000°C (C) 5000°C (D) 1000°C
- Equal masses of methane and oxygen are mixed in an empty container at 25°C . The fraction of total pressure exerted by oxygen is:
(A) $\frac{1}{3}$ (B) $\frac{8}{9}$ (C) $\frac{1}{9}$ (D) $\frac{16}{17}$

MCQs Ans Key.

Q:1 (B)

Q:2 (C)

Q:3 (B)

Q:4 (A)

Q:5 (C)

Q:6 (D)

Q:7 (D)

Q:8 (C)

Q:9 (D)

Q:10 (B)

Q:11 (D)

Q:12 (B)

Q:13 (B)

Q:14 (B)

Q:15 (C)

Q:16 (A)

Q:17 (B)

Q:18 (C)

Q:19 (A)

Q:20 (A)